

Review:

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2. What do elements in the same group have in common? (at least 2 things)

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Energy Levels

2. What do elements in the same group have in common? (at least 2 things)

- Valence e⁻
- Similar properties
- Same charge

Review:

3. What happens to the atomic radius as you go down a group?
4. Which group on the periodic table is more reactive: alkali metals or alkaline earth metals? Why?
5. Elements gain or lose electrons to form.....

Review:

3. What happens to the atomic radius as you go down a group?

increases, b/c you gain energy levels

4. Which group on the periodic table is more reactive, alkali metals or alkaline earth metals? Why?

b/c they're closer to an octet
(lose $1e^-$ vs. $2e^-$)

5. Elements gain or lose electrons to form.....

an octet!

Review:

6. What do halogens do when they form ions? What kind of charge do they have?

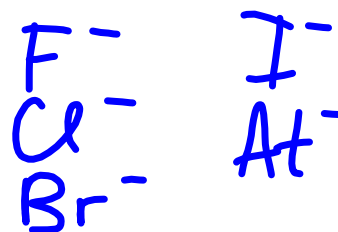
7. What would be the charge of a Ca ion?

Review:

6. What do halogens do when they form ions? What kind of charge do they have?

gain e^-

$$0 + (-1) = \boxed{-1}$$



7. What would be the charge of a Ca ion?

$$0 - (-2) = +2 \quad (\text{Ca}^{2+})$$

Review:

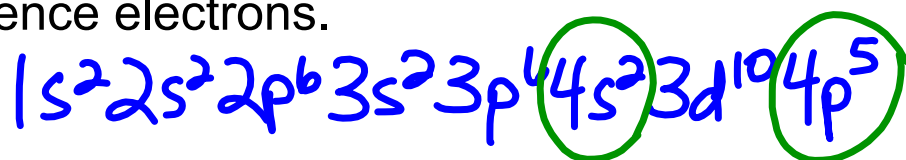
8. Write the electron configuration for Br, circle the valence electrons.

9. Put the following in order of increasing electronegativity: Na, Al, Rb

10. Put the following in order of decreasing atomic radius: Na, Al, Rb

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