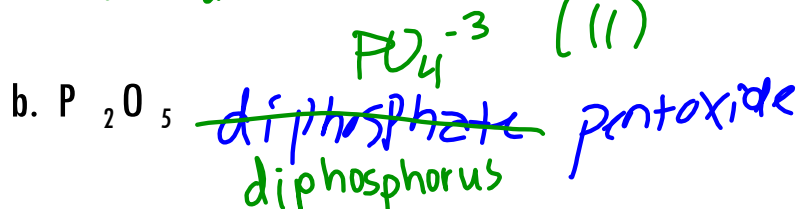


Practice:

Name the following compounds:



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Practice:

Write the following compounds:

a. nitrogen monoxide



b. magnesium hydroxide



c. tricarbon octahydride



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Practice:

Classify the compounds as ionic, non-polar covalent, polar covalent:

a. Cs-F

.79 3.98 ionic

b. N-S

3.04 2.58 polar

c. C-F

2.55 3.98

polar covalent



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Practice:

Draw the Lewis structure:

a. HCN

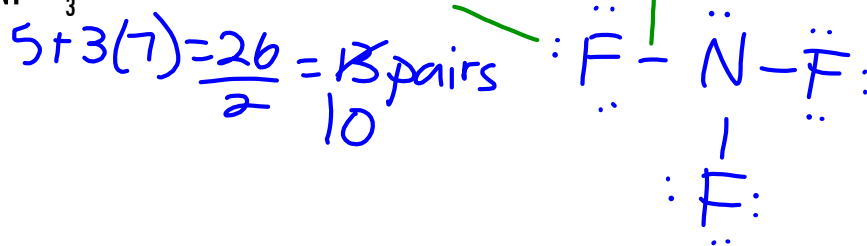
$$1 + 4 + 5 = \frac{10}{2} = 5 \text{ pairs}$$

have
need

$$\begin{array}{c} 2 \times 0 \\ \hline 5 \times 0 \end{array}$$



b. NF₃



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Practice:

Determine if the following describes an ionic or covalent compound:

1. Can be nonpolar due to equal electron sharing. ~~I~~ C
2. Formula must have balanced charges. ~~I~~ I
3. Formed by oppositely charged ions attracting. ~~I~~ I
4. Have very high melting points and boiling points. ~~I~~ I
5. Formed between a metal and a nonmetal. ~~I~~ I
6. Has bonding and unshared electron pairs. ~~I~~ C
7. Formed by a shared pair of electrons. C

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