

Daily Question 9/15

1. Which of the following atoms are isotopes of each other?
 - a. $p=6, n=6, e=6$ and $p=6, n=7, e=6$
 - b. $p=13, n=14, e=13$ and $p=14, n=14, e=13$
 - c. $p=8, n=8, e=8$ and $p=8, n=7, e=8$
 - d. $p=10, n=11, e=11$ and $p=11, n=10, e=11$
2. Write the following isotope 2 different ways, $p=19, n=21, e=19$.
3. How many electrons will a neutral atom of magnesium have?

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d. p=10, n=11, e=11 and p=11, n=10, e=11

2. Write the following isotope 2 different ways, p=19, n=21, e=19.

potassium-40

$^{40}_{19}\text{K}$

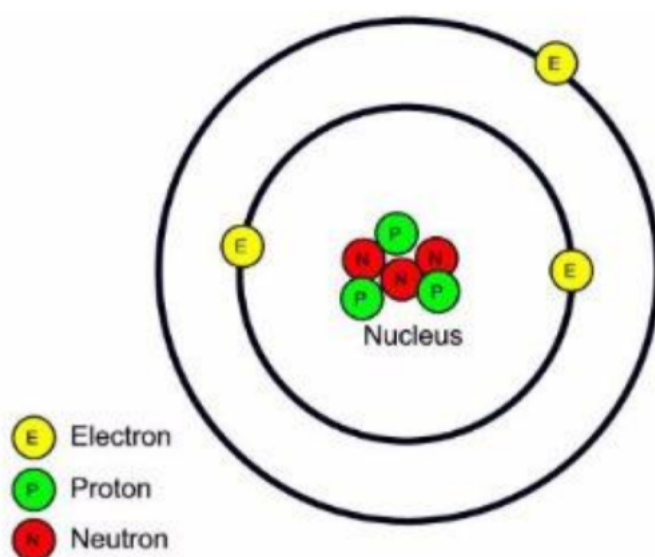
3. How many electrons will a neutral atom of magnesium have?

12 electrons

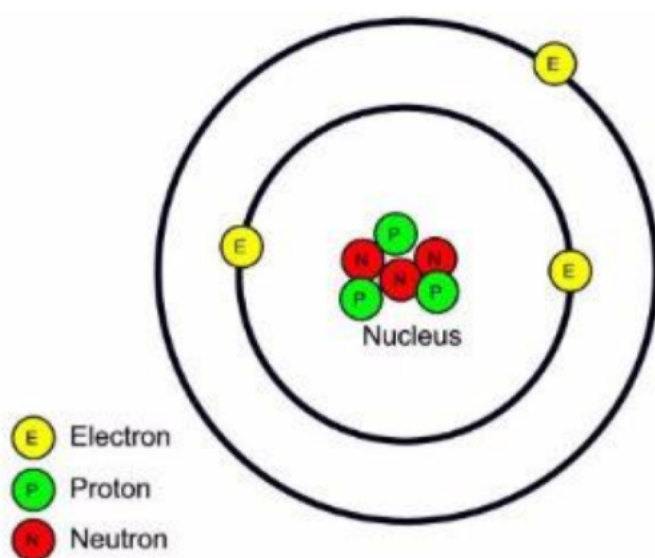
4. Find the average atomic mass for Mg if 78.99% of Mg atoms are magnesium-24, 10.00% are magnesium-25 and 11.01% are magnesium-26.

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1. What atom is represented in the atom shown below?

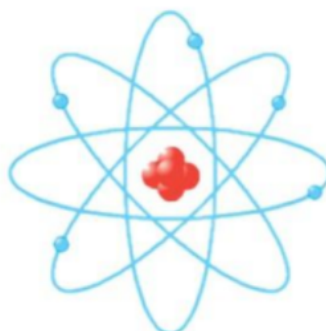
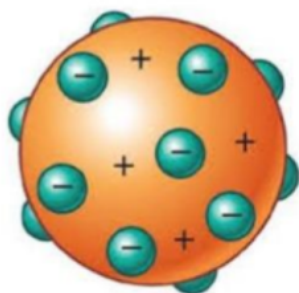


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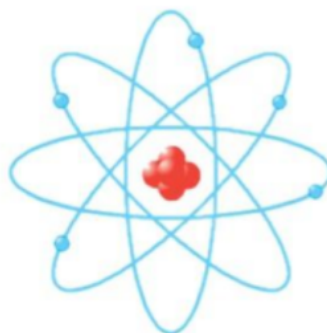
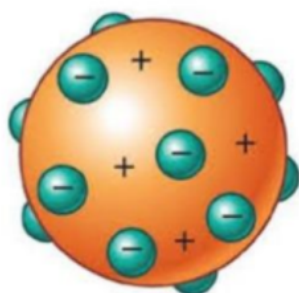


Lithium

2. What is the difference between the two atoms below? Which one is more accurate?



2. What is the difference between the two atoms below? Which one is more accurate?



1. the one on the right has a nucleus and the one on the left doesn't

2. the one on the right is more accurate

**What is the mass of an atom containing 10 protons,
10 electrons and 11 neutrons?**

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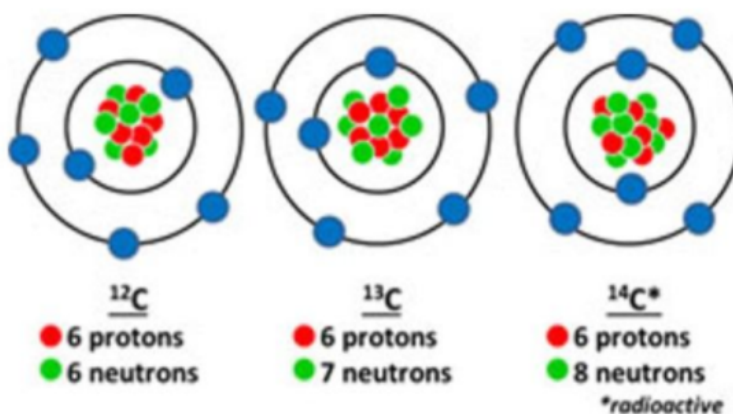
21 amu

How many electrons will a neutral atom of phosphorus have?

How many electrons will a neutral atom of phosphorus have?

15 electrons

Write the long isotope notation for the isotopes below.



carbon-12; carbon-13; carbon-14